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MHT CET 2021 Marking Scheme.

Candidates appearing for the exam must know the marking scheme. As per the MHT CET 2021 exam pattern in Paper, I (Mathematics) candidates will be allotted two marks for each question, whereas in paper II (Physics and Chemistry), one mark will be given for

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each correct answer. There will be no negative marking, and for unattempted answers, no marks will be deducted.

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offered by both JEE Main exam conducting bodies and various private institutions in order to get candidates prepared for JEE Main.

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Answer (1 of 166): The best book for organic chemistry are 1.Ncert and Ncert

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exemplar. 2. Organic Chemistry –
Himanshu Pandey (GRB Publication).

Which are the best books for organic chemistry for ... - Quora

Chemistry:

Atkins_Physical_Chemistry_8th_Edition;

Concise Inorganic Chemistry (4th

Edition) by J.D.Lee; Organic Chemistry

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Himanshu Pandey; Organic-Chemistry-8
E-2013-L-G-Wade-Solution-Manual;
Physical Chemistry - Narendra Avasthi-1;
Physical-Chemistry-by-P-Bahadur; RC
Mukherjee - Modern Chemical Calc;
Mathematics. RD Sharma part 1 JEE
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Organic Chemistry by Solomons and Fryhle for Organic Chemistry; Concise Inorganic Chemistry by J. D. Lee for Inorganic Chemistry; Modern Approach

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Molecular Similarity. Molecular similarity is one of the most heavily exploited concepts in cheminformatics and related areas (such as medicinal chemistry and drug discovery). It is applied to multiple tasks, including similarity searching , property prediction , synthesis design , virtual screening [2,3,6], cluster analysis [7,8], and molecular diversity analysis [1].

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6.1: Molecular Descriptors - Chemistry LibreTexts

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reduction in syllabus by CBSE and other state boards.

JEE Main Syllabus 2022 for B. E/ B.tech, B ... - Collegedunia

Numerical problems based On Mole
Concept Question 1. Calculate the mass
of 6.022×10^{23} molecule of Calcium
carbonate (CaCO_3). Solution — Molar

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mass (Molecular mass in gram) of
 $\text{CaCO}_3 = 40 + 12 + 3 \times 16 = 100 \text{ g}$ No. of
moles of $\text{CaCO}_3 = \frac{\text{No. of molecules}}{\text{Avogadro constant}} = \frac{6.022 \times 10^{23}}{6.022 \times 10^{23}} = 1 \text{ mole...}$

**Problems Based On Mole Concept
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